PREPARATION OF Cu MATRIX COMPOSITE REINFORCED WITH IN-SITU
NANOSIZED Al₂O₃ PARTICLE POWDER FROM METAL NITRATES

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Abstract
The objective of the present work is to investigate the feasibility of the synthesis of copper matrix
composite reinforced with in-situ nanosized Al₂O₃ particle powder via combustion synthesis
method from metal nitrates followed by reducing process at high temperature. The starting nitrates
Cu(NO₃)₂·3H₂O and Al(NO₃)₃·9H₂O composition corresponds to Cu-30%Al₂O₃. X-ray
Diffraction (XRD) patterns of the obtained powders indicated the presence of the oxides CuO and
CuAl₂O₄. The powder had the size of 75 ± 10 nm after deagglomerating by soft ball milling for
24h. After reducing in CO at 1000°C for 3h, the peaks of the oxides were no longer observed and
were replaced by the peaks of Cu and α-Al₂O₃. The morphology of the reduced powders observed
by scanning electron microscopy (SEM) and energy dispersive X-ray spectroscopy (EDS) analysis
showed well distribution of the α-Al₂O₃ particles within the Cu matrix with an average particle
size of 40 nm.

Keywords: Cu-Al₂O₃ nanocomposite, CuAl₂O₄, CuAlO₂, combustion synthesis, reduction

1 Introduction
The aim of fabricating copper-based composites reinforced with dispersed ceramic particles is to
enhance the mechanical properties of copper, in particular their higher specific strength and good
elevated temperature mechanical properties while still maintaining the high electrical and thermal
conductivities [1-8]. Currently, there are two main routes of producing such kind of composites,
which are ex-situ and in-situ processes. In in-situ synthesis technique, the reinforcing ceramic
phases are synthesized in the metallic matrix by a chemical reaction during the composite
fabrication. Consequently, compared to ex-situ synthesis technique, in-situ routes introduce
significant advantages such as nanosize, good distribution and thermodynamical stability of
reinforcing particulate phases [9-13], which considerably enhance the properties of the
composites.

The Cu nanocomposite reinforced by in-situ Al₂O₃ have been developed through different
synthesis routes such as mechanical alloying and rapid solidification. Ying and Zhang studied the
synthesis of a Cu-20 vol.% Al₂O₃ nanocomposite via mechanical milling of a Cu-Al powder
together with CuO powder [14]. Al₂O₃ particles in the consolidated composite material have a size
smaller than 200 nm in diameter. Recently, chemical route has emerged as a newly developed
method to prepare Cu nanocomposites reinforced with nanosized Al₂O₃ particles. Cu
nanocomposites reinforced by 5, 10, and 15 wt.% Al$_2$O$_3$ were prepared using mechano-chemical method [15-18]. This research reported that Cu was added to aqueous solution of aluminum nitrate or aqueous solution of aluminum nitrate and ammonium hydroxide. The average particle size of Cu and Al$_2$O$_3$ were 209 nm and 50 nm, respectively. Krakum et al. prepared Cu-Al$_2$O$_3$ composite via directly mixing of CuO and Al$_2$O$_3$ powders [19]. The powder mixture was then ball-milled in ethanol and sintered in an SPS apparatus to achieve highly dense CuAlO$_2$ sample. This bulk sample was then reduced by H$_2$ gas to obtain Cu-Al$_2$O$_3$ composite. A new process to produce homogeneous Cu-Al$_2$O$_3$ nanocomposites from combustion reaction of metal nitrates was developed in the present study. The final Cu-Al$_2$O$_3$ nanocomposites was obtained by CO reducing and sintering process. The phases present, morphology and microstructure will be determined by X-ray Diffraction (XRD), Scanning electron microscopy (SEM) and Energy dispersive X-ray spectroscopy analysis (EDS) techniques.

2 Experimental procedure
Copper nitrate Cu(NO$_3$)$_2$.3H$_2$O (>99%) and aluminum nitrate (Al(NO$_3$)$_3$.9H$_2$O (>99%) as oxidizers and urea CH$_4$N$_2$O (>99%), as fuel were used to produce Cu-30 vol% Al$_2$O$_3$ nanocomposite, purchased from Xilong Scientific Co. Ltd., China. Precursor materials were mixed by a fixed ratio, and then dissolved in distilled water. Subsequently, the solution was placed in an electric resistance furnace (Linn HT1300, Germany) which was preheated at 500°C. The combustion reaction was performed according to the reactions (1) and (2).

$$3\text{Cu(NO}_3\text{)}_2\text{(aq)} + 6\text{Al(NO}_3\text{)}_3\text{(aq)} + 20\text{CH}_4\text{N}_2\text{O (aq)} \rightarrow 3\text{CuAl}_2\text{O}_4\text{(s)} + 20\text{CO}_2\text{(g)} + 40\text{H}_2\text{O(g)} + 32\text{N}_2\text{(g)} \quad (1.)$$

$$3\text{Cu(NO}_3\text{)}_2\text{(aq)} + 5\text{CH}_4\text{N}_2\text{O (aq)} \rightarrow 3\text{CuO(s)} + 5\text{CO}_2\text{(g)} + 10\text{H}_2\text{O(g)} + 8\text{N}_2\text{(g)} \quad (2.)$$

The synthesized product was milled for 24 hours in a pure ethanol solution (96%) using alumina balls with ball-to-powder mass ratio of 20/1 then dried at 120°C for 24 hours and calcined at 1100°C for 2 hours. The reduction process was carried out at 1000°C for 3 hours by annealing in a forming gas (99.97% CO) environment. The green compacts were formed by uniaxial pressing in a 10 mm inner diameter cylindrical steel die under a uniaxial applied pressure of 500 MPa. The compacted samples were then sintered by a tube furnace (Linn HT1200, Germany) at 1000°C for 3 hours with a heating rate of 5°C/min in argon atmosphere. The phase analysis was carried out by XRD (D5000 Siemens, Germany) using Cu K$\alpha$ radiation. The morphology of the synthesized powders and the sintered samples was characterized by a field-emitting scanning electron microscope (FE-SEM Hitachi S4800, Japan). The particle size and the size distribution were evaluated by ImageJ software through SEM images. In addition, EDS was performed to identify the elements that present in the synthesized powders.

3 Results and Discussion

![Graph](image-url)
The XRD patterns of (a) the combustion-synthesized powder, (b) after annealed at 1100°C in air for 2 hours and (c) after reduction in CO at 1000°C for 3 hours

**Fig. 1** The XRD patterns of (a) the combustion-synthesized powder, (b) after annealed at 1100°C in air for 2 hours and (c) after reduction in CO at 1000°C for 3 hours

The evolution of the microstructure organization of the combustion-synthesized powder according to each step of the thermal treatment process can be deduced from XRD spectra (**Fig. 1**). The XRD spectrum of the combustion-synthesized powder exhibited almost all the main CuO and CuAl₂O₄

**Fig. 2** SEM images of combustion-synthesized powder annealed at 1100°C in air for 2 hours in (a) SE mode, (b) BSE mode and its corresponding EDS patterns acquired on (c) A and (d) B areas

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reflections (Fig. 1a). The appeared peaks were very broad indicating the poor crystalline nature of combustion synthesized product due to nanosized effect. After 2-hour annealing at 1100°C in air, the peaks of CuO phase was more intense and well-defined indicating a good crystallinity of the annealed product. However, the CuAl2O4 peaks disappeared and replaced by the peaks of CuAlO2 (Fig. 1b). According to Hu et al., [20] the decomposition of CuAl2O4 at a temperature higher than 900°C led to the formation of CuAlO2 and α-Al2O3 or γ-Al2O3 phases depending on the partial pressure of O2. However, if the amount of CuO is excessive, the solid–solid reaction between CuO and CuAl2O4 can take place to form CuAlO2. The XRD pattern of annealed powder after reduction in the presence of reducing gas CO at 1000°C for 3 hours showed the sharp and high intense peaks of metallic Cu (Fig. 1c). The small peak width indicated a good crystallinity with a large grain size of the formed Cu. However, no peaks of Al2O3 phase was observed in XRD spectrum. The invisibility of Al2O3 could be due to several reasons including poor crystallinity and small particle size of the generated Al2O3 and its much lower backscattering factor compared to Cu.

The SEM image in secondary electron mode of the combustion-synthesized powder after annealing was displayed in Fig. 2 showing a monomodal particle size distribution with an average size of 100 nm. Back-scattered SEM images of the annealed powder illustrated the contrast between CuO and CuAlO2 (Fig. 2b). EDS analysis (Fig. 2c,d) indicates that the bright area (marked as A) and the dark area (marked as B) are CuAlO2 and CuO, respectively.

![SEM images](image)

Fig. 3  SEM images of the composite powder reduced by CO at 1000°C for 3 hours in (a) powder, (b) its polished surface and corresponding EDS patterns acquired on (c) dark and (d) white areas

The reduction in the presence of reducing gas CO at 1000°C for 3 hours led to the formation of pores, which were uniformly distributed along the grain surfaces (Fig. 3a). In order to investigate
the distribution homogeneity of reinforcing particulates, back-scattered SEM (mode) and corresponding EDS of the composite powder are presented in Fig. 3b–d. The EDS spectra which acquired on the dark (Fig. 3d) and white (Fig. 3c) regions clearly demonstrated the existences of Cu matrix and Al₂O₃ particulate reinforcement. In the composites, the in-situ Al₂O₃ particles with an average size of 70 nm were uniformly dispersed throughout in the Cu matrix, as seen in Fig. 3b.

**Fig. 4** XRD patterns of the Cu-30 vol% Al₂O₃ composite sintered at 1000°C for 3 hours

The XRD pattern of the sintered Cu-30 vol% Al₂O₃ composite is plotted on a logarithmic scale (Fig. 4) in order to observe the low-intense peaks. Apart from the peaks of copper matrix, the
XRD pattern shows the reflections corresponding to the hexagonal close-packed structure of α-Al₂O₃. However, these Al₂O₃ peaks were not observed after reducing in CO. Their appearance could be due to that the heat treatment recovered considerably internal strains and made the crystallite size increase remarkably. SEM images of the sintered composite in both (a) secondary electron and (b) back-scattered electron modes from the same area and their corresponding EDS patterns acquired on (c) bright and (d) dark areas (Fig. 5a-d) proved a nanosized and well-dispersed Al₂O₃ particles within Cu matrix. The average size of the in-situ reinforced Al₂O₃ particles was about 40 nm. In the future works, the fabrication of bulk nanocomposite with different reinforcement composition and their mechanical and physical properties will be studied.

4 Conclusions
Cu-30vol.% Al₂O₃ nanocomposites have been successfully synthesized by chemical routes through combustion reaction of metal nitrates (i.e. Al(NO₃)₃ and Cu(NO₃)₂) followed by CO reduction and sintering. The obtained XRD spectra revealed that the mixture of CuO and CuAl₂O₄ powder formed after combustion reaction was transferred to CuO and CuAlO₂ after annealing and to composite powder of Cu/Al₂O₃ after reducing by CO. SEM images in both SE and BSE mode and its corresponding EDS patterns showed nanosized and well-dispersed of Al₂O₃ particles within Cu matrix. The average size of the in-situ reinforced Al₂O₃ particles was 40 nm.

References

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